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Disassembling Water, H₂O Electrolysis

The purpose of this project is to see what electrolyte is best for Electrolysis of water to occur. The hypothesis is that a solution of salt and distilled water will act as a better electrolyte than a solution of Epsom salts and distilled water. First, construct electrodes using syringes and graphite rods from batteries. Then, fasten alligator clips to wires with a soldering iron. Mix solutions of salt and distilled water, Epsom Salts and distilled water, and have a control of just distilled water. Connect the electrodes to batteries. There was no activity on either wire from the container of just distilled water. There was activity on the red wire of each other solution, but it was minor. The Sodium Chloride solution had twice as much activity at first and eventually four times the amount as the Epsom Salts. In conclusion, my hypothesis was proved correct, the solution containing Sodium Chloride produced more Hydrogen and Oxygen, thus showing that table salt was a better electrolyte.